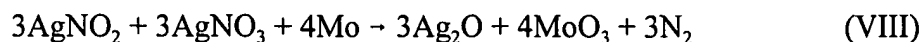


In the Specification:

Amend as follows at page 18:

Example 3.



C/ A comelt of equimolar amounts of  $\text{AgNO}_2$  and  $\text{AgNO}_3$  was mixed with a stoichiometric amount of Mo metal in accordance with equation VIII, i.e., 34.1% by weight  $\text{AgNO}_2$ , 37.6% by weight  $\text{AgNO}_3$ , and 28.3% by weight Mo. An autoignition temperature of  $131 \pm 2^\circ\text{C}$  was determined for the composition using DSC.

✓ Example 4.



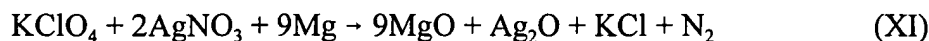
~~————— Lithium perchlorate,  $\text{LiClO}_4$ , was mixed with a stoichiometric amount of Mo in accordance with equation IX, i.e., 45.4% by weight  $\text{LiClO}_4$  and 54.6% by weight Mo. An autoignition temperature of  $147 \pm 2^\circ\text{C}$  was determined for the composition using DSC.~~

C<sup>2</sup> Example 5.



$\text{AgNO}_3$  was mixed with a stoichiometric amount of magnesium, Mg, metal in accordance with equation X, i.e., 73.7% by weight  $\text{AgNO}_3$  and 26.3% by weight Mg. An autoignition temperature of  $157 \pm 2^\circ\text{C}$  was determined for the composition using DSC.

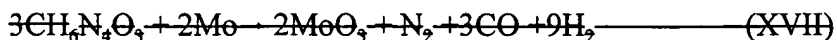
C<sup>3</sup> Example 6.



Amend as follows at page 21:

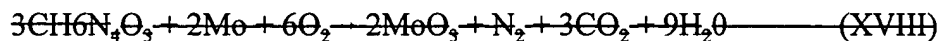
CH  
metal in accordance with equation XVI, i.e., 20.5% by weight  $\text{NaNO}_3$ , 41.0% by weight  $\text{AgNO}_3$  and 38.5% by weight Mo. The composition autoignited at  $217 \pm 2^\circ\text{C}$  by DSC analysis.

✓  
Example 12:



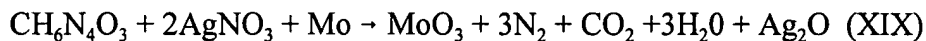
~~\_\_\_\_\_ Guanidine nitrate,  $\text{CH}_6\text{N}_4\text{O}_3$ , was mixed with a stoichiometric amount of Mo in accordance with equation XVII, i.e., 60.4% by weight  $\text{CH}_6\text{N}_4\text{O}_3$  and 39.6% by weight Mo. The composition autoignited at  $230 \pm 2^\circ\text{C}$  by DSC analysis.~~

~~\_\_\_\_\_ This is an underoxidized reaction which leaves some products in an incompletely oxidized state. If there is an external source of oxygen the reaction proceeds according to equation XVIII.~~



~~\_\_\_\_\_ This composition points out the utility of using organic nitrates in autoignition reactions.~~

Example 13.



A 1:2 ratio of guanidine nitrate to  $\text{AgNO}_3$  was mixed with a stoichiometric amount of Mo in accordance with equation XIX, i.e., 21.9% by weight  $\text{CH}_6\text{N}_4\text{O}_3$ , 60.9%  $\text{AgNO}_3$  and 17.2% by weight Mo. The composition autoignited at  $172 \pm 2^\circ\text{C}$  (by DSC).

Q15  
Concl'd.

This composition is also an example of organic nitrates in autoignition reactions. However, this composition is fully oxidized, and, therefore, requires no external source of oxygen.

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